

**GENERAL CERTIFICATE OF SECONDARY EDUCATION
GATEWAY SCIENCE
SCIENCE B**

Unit 2 Modules B2 C2 P2
(Foundation Tier)

B622/01

Candidates answer on the question paper
A calculator may be used for this paper

OCR Supplied Materials:

- None

Other Materials Required:

- Pencil
- Ruler (cm/mm)

**Friday 12 June 2009
Morning**

Duration: 1 hour



Candidate Forename		Candidate Surname	
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Centre Number						Candidate Number			
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INSTRUCTIONS TO CANDIDATES

- Write your name clearly in capital letters, your Centre Number and Candidate Number in the boxes above.
- Use black ink. Pencil may be used for graphs and diagrams only.
- Read each question carefully and make sure that you know what you have to do before starting your answer.
- Answer **all** the questions.
- Do **not** write in the bar codes.
- Write your answer to each question in the space provided, however additional paper may be used if necessary.

INFORMATION FOR CANDIDATES

- The number of marks is given in brackets [] at the end of each question or part question.
- A list of physics equations is printed on page two.
- The Periodic Table is printed on the back page.
- The total number of marks for this paper is **60**.
- This document consists of **20** pages. Any blank pages are indicated.

EQUATIONS

$$\text{efficiency} = \frac{\text{useful energy output}}{\text{total energy input}}$$

$$\text{wave speed} = \text{frequency} \times \text{wavelength}$$

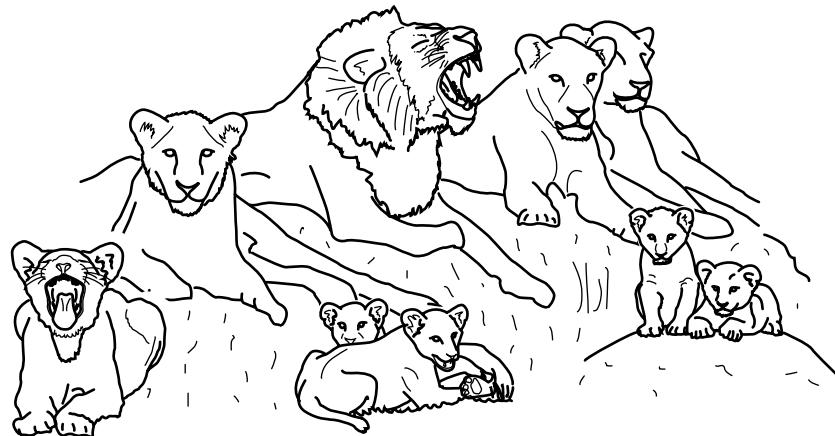
$$\text{power} = \text{voltage} \times \text{current}$$

$$\text{energy (kilowatt hours)} = \text{power (kW)} \times \text{time (h)}$$

Answer **all** the questions.

Section A – Module B2

- 1 The picture shows a group of lions.



- (a) The lions show variation.

Describe **two** ways you can see in the picture that they show variation.

1

2 [2]

- (b) Look at the list.

parasite

predator

prey

Which word best describes a lion?

Choose **one** word from the list.

..... [1]

- (c) Describe and explain **one** way that lions are adapted to survive.

How they are adapted

How this helps them to survive

..... [2]

[Total: 5]

- 2 Complete the sentences about photosynthesis.

During photosynthesis, plants take in a gas called from the air.

They also take in a liquid called from the soil.

The type of food that plants make in photosynthesis is

During photosynthesis, plants make a gas called [4]

[Total: 4]

- 3 (a) Look at the list of animals in the table.

Put ticks (✓) in the table to show whether each animal is **extinct** or **endangered**.

animal	is it extinct?	is it endangered?
dodo		
mammoth		
panda		
sabre-toothed tiger		

[2]

- (b) What do the terms **extinct** and **endangered** mean?

Extinct means

.....

Endangered means

.....

[2]

- (c) The increasing human population has caused some animals to become extinct.

Suggest **one** reason why.

.....

[1]

[Total: 5]

- 4 Lynne is investigating some of the animals and plants in a wood.



- (a) Lynne notices that small bushes grow in some of the spaces between the trees, but **not** under the trees.

Suggest why small bushes do **not** grow under the trees.

.....
.....

[1]

- (b) Lynne is investigating peppered moths in the wood.

Some peppered moths are pale. Some are dark.



pale peppered moth

dark peppered moth

Lynne counts the number of both types of peppered moths on ten trees.

The table shows her results.

tree number	number of pale peppered moths	number of dark peppered moths
1	1	0
2	0	1
3	1	0
4	3	0
5	0	1
6	1	0
7	0	0
8	0	0
9	2	0
10	0	0

- (i) Lynne notices that there are more pale peppered moths than dark peppered moths.

She knows that there are 300 trees in the wood.

Lynne uses this information to estimate that there are 60 dark peppered moths in the whole wood.

Use the information given to estimate the number of pale peppered moths in the whole wood.

You are advised to show your working.

answer

[2]

- (ii) Suggest **one** reason why there are more pale peppered moths than dark peppered moths in the wood.

..... [1]

- (iii) The two types of peppered moths both belong to the same species.

How could Lynne show this?

.....
.....
..... [2]

[Total: 6]

Section B – Module C2

- 5 Pete and Sally investigate marble and limestone.
Limestone and marble both have the formula, CaCO_3 .

- (a) What is the chemical **name** for limestone and marble?

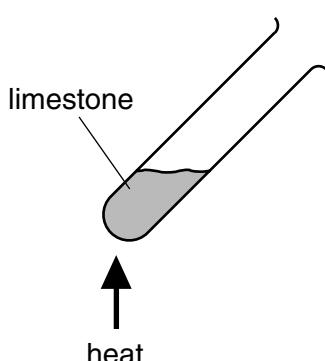
..... [1]

- (b) Limestone and marble are both building materials.

Write down the name of **one other** building material.

..... [1]

- (c) Sally heats some limestone.



- (i) A gas is made.

Write down the name of this gas.

..... [1]

- (ii) When limestone is heated, **thermal decomposition** happens.

What is thermal decomposition?

..... [1]

- (d) Limestone is used to make cement.

Limestone is mixed with another substance.

Write down the name of this substance.

Choose from the list.

clay
glass
granite
iron ore

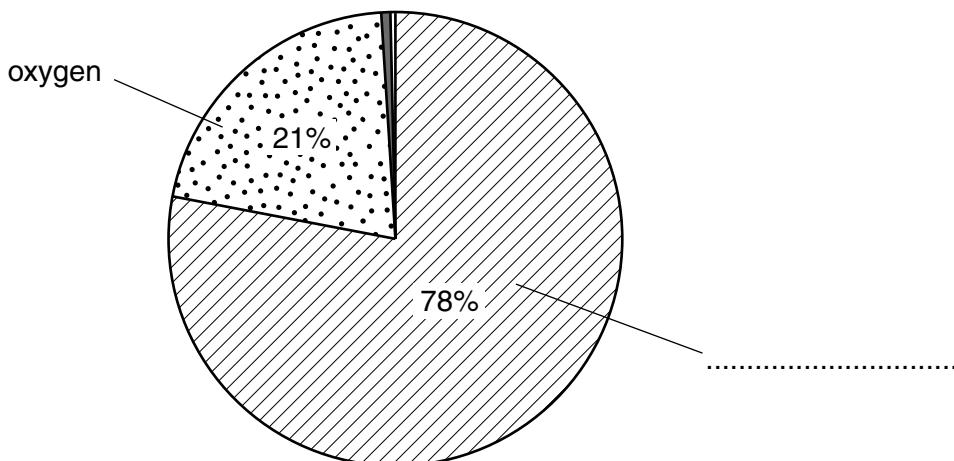
answer [1]

[Total: 5]

Turn over

- 6 This question is about gases in the air.

Look at the pie chart. It shows the composition of the air.



- (a) Complete the pie chart to show which gas makes up **78%** of the air.

[1]

- (b) Sulfur dioxide causes air pollution.

Write about sulfur dioxide pollution.

Your answer should include

- what is made during sulfur dioxide pollution
- two effects of sulfur dioxide pollution.

.....
.....
.....
.....

[3]

- (c) Carbon monoxide also causes air pollution.

It is made when petrol burns in a car engine.

Carbon monoxide is removed from car exhaust gases.

What is the name of the equipment which removes carbon monoxide?

.....

[Total: 5]

- 7 This question is about paints.



- (a) Paint is used to paint the front door of John's house.

Write down **one** reason why John paints his front door.

..... [1]

- (b) Paints are made up of

binding medium

colouring

solvent

Which one thins the paint and makes it easier to use?

Choose from the list.

answer [1]

- (c) Some pigments used in paint change colour when they are heated.

They are called **thermochromic pigments**.

Write down **one** use of thermochromic pigments.

..... [1]

[Total: 3]

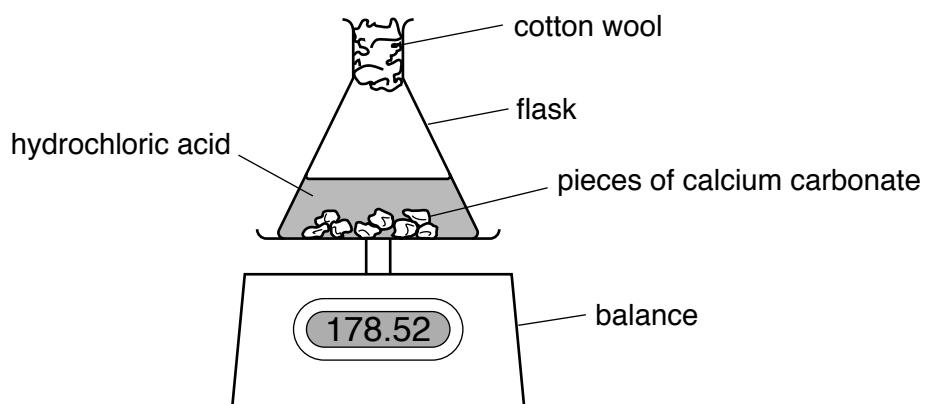
- 8 Fred and Sue investigate the reaction of pieces of calcium carbonate and hydrochloric acid.
- Carbon dioxide is given off during the reaction.
- Calcium chloride and water are also made.
- (a) Write a **word** equation for this reaction.

..... [1]

- (b) Fred and Sue measure the mass of the reaction mixture every 30 seconds during the experiment.

Look at the diagram.

It shows the apparatus they use.



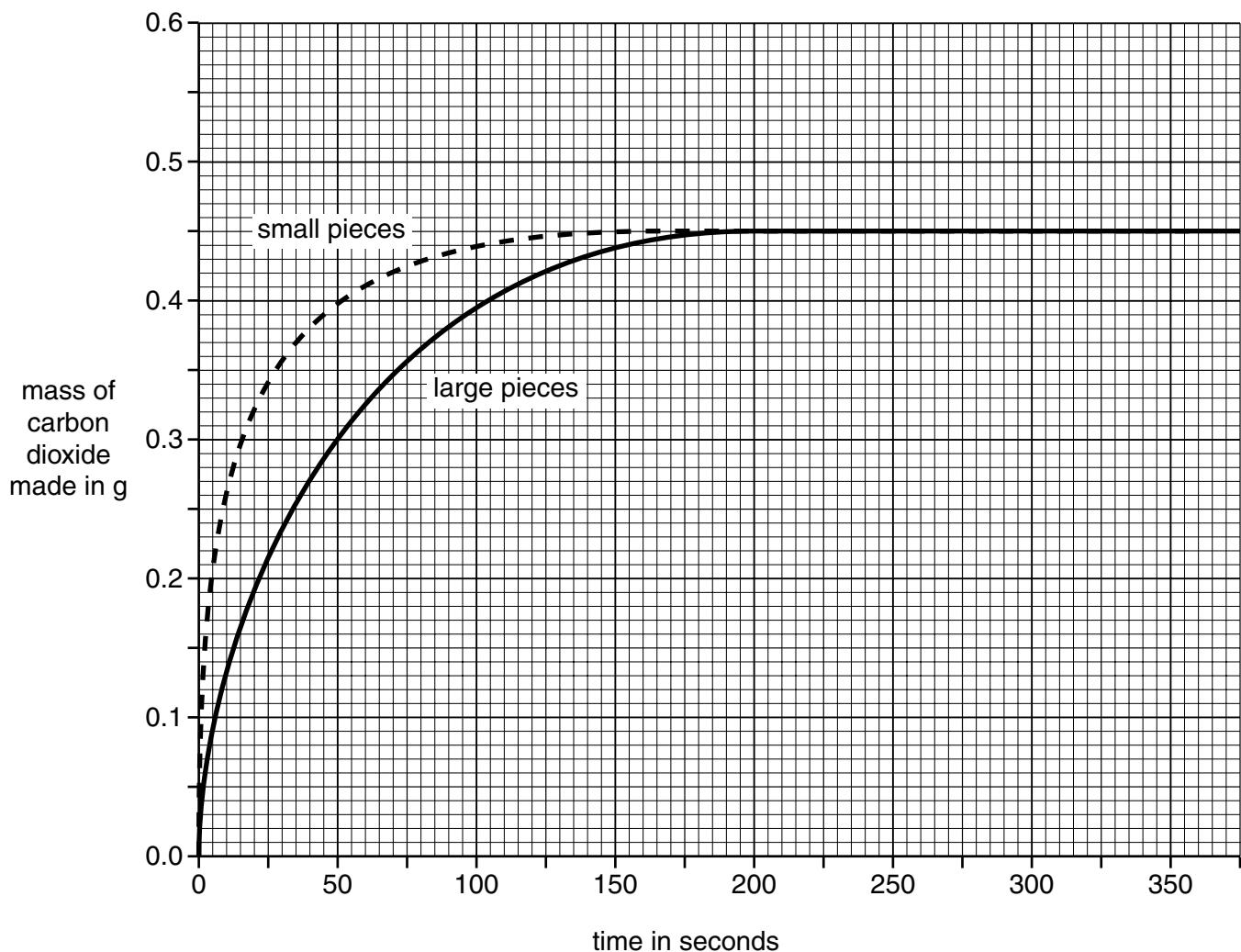
After every measurement, Sue works out the total mass of carbon dioxide given off.

They do the experiment again.

They use the same amounts of acid and calcium carbonate.

This time they use **smaller** pieces of calcium carbonate.

Look at the graph. It shows their results.



- (i) Look at the curve for the **small** pieces.

What mass of carbon dioxide is given off after 50 seconds?

..... g

[1]

- (ii) Look at the curve for the **large** pieces.

How long does it take for this reaction to finish?

..... seconds

[1]

- (c) Calcium carbonate is left in the flask at the end of both experiments.

Why do both reactions stop?

.....

[1]

- (d) The reaction using small pieces is faster than the reaction using large pieces.

Explain why. Use ideas about particles.

.....

[1]

[Total: 5]

Turn over

9 Look at the formulas.



(a) Which one of the formulas contains three oxygen atoms only?

..... [1]

(b) Which one of the formulas contains a **total** of six atoms?

..... [1]

[Total: 2]

Section C – Module P2

10 Look at the list of objects.

Draw a line between each **object** and its correct **description**.

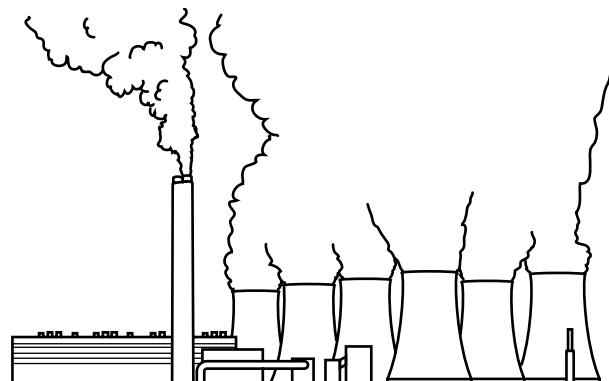
One has been done for you.

object	description
asteroid	light cannot escape from it
black hole	has a tail made of water vapour and debris
comet	a planet that orbits the Sun
Earth	transfers light into electricity
photocell	transfers energy to Earth as light and heat
Sun	a rock in space

[4]

[Total: 4]

11 Electricity is produced by power stations.



(a) Many power stations burn fuels to get electricity.

(i) Name one **fossil** fuel burnt in power stations.

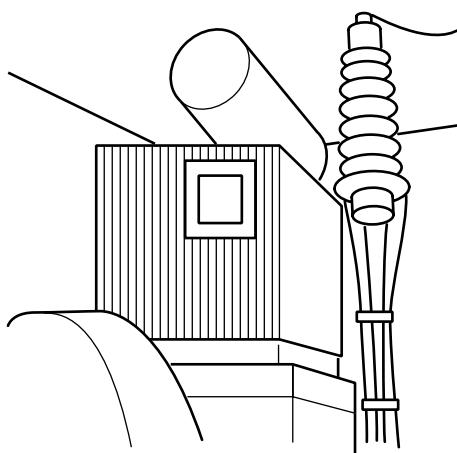
..... [1]

(ii) Name one **renewable** fuel burnt in power stations.

..... [1]

(b) Electricity leaves a power station through a transformer.

Look at the diagram of a transformer.



(i) What does a transformer do?

..... [1]

(ii) The electricity is sent to consumers.

Write down **one** example of a consumer of electricity.

..... [1]

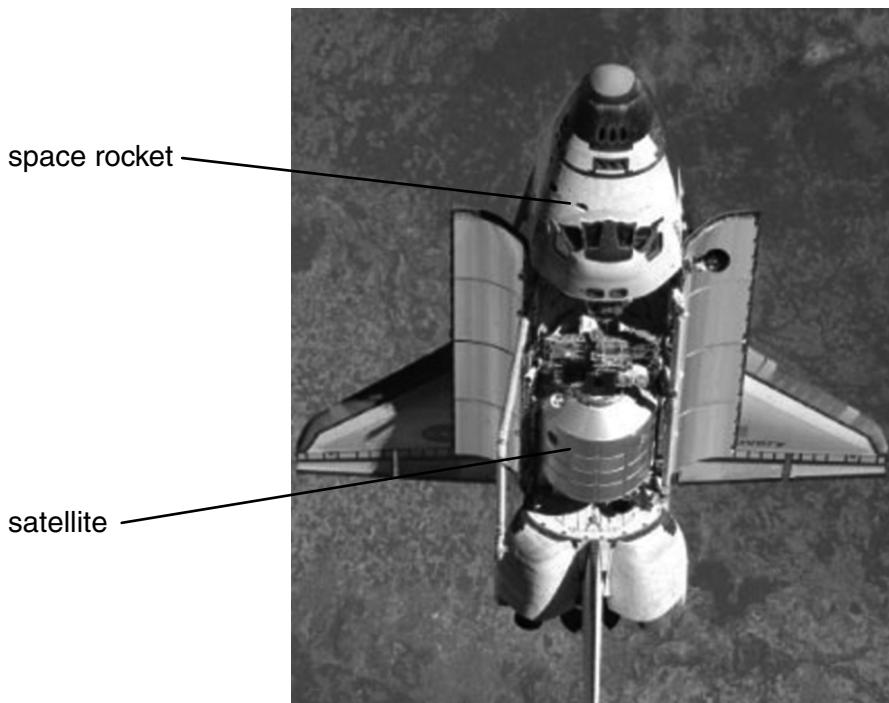
(iii) How does the electricity get to the consumers?

..... [1]

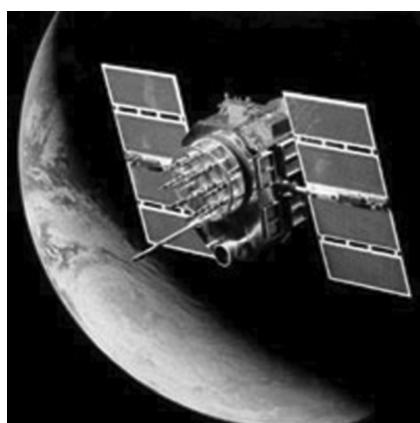
[Total: 5]

12 Space rockets can carry people (astronauts) into space.

(a) Some rockets carry satellites into space.



The **satellite** is released and it orbits the Earth.

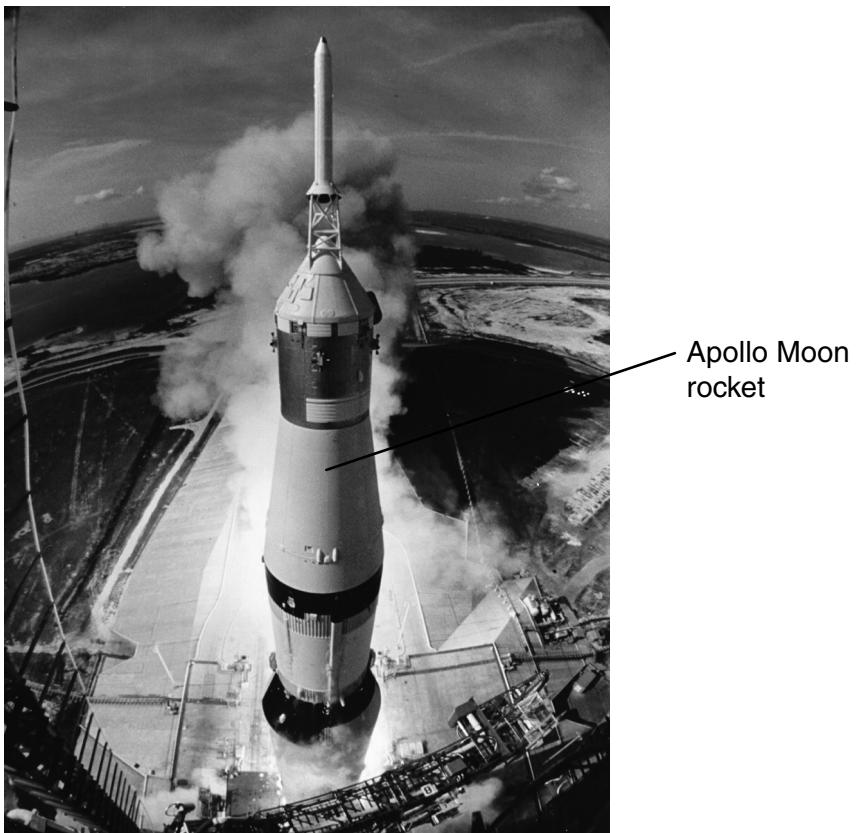


Write down two **uses** of satellites.

1

2 [2]

- (b) Astronauts use space rockets to visit the Moon.



After this journey they return to Earth.

This takes several days.

The astronauts need things to keep them alive.

Write about what they need to keep them alive.

.....

.....

.....

.....

[3]

[Total: 5]

13 This question is about nuclear radiation.

- (a) The three types of nuclear radiation are alpha, beta and gamma.

They can all be used in cancer treatment.

- (i) Write down one other use of **alpha** radiation.

..... [1]

- (ii) Write down one other use of **beta** radiation.

..... [1]

- (iii) Write down one other use of **gamma** radiation.

..... [1]

- (b) Background radiation is around us all the time.

Write down one source of this background radiation.

..... [1]

- (c) A nuclear power station uses uranium as a fuel.

- (i) Why do we get **plutonium** in this nuclear power station?

..... [1]

- (ii) What is **plutonium** used for?

..... [1]

[Total: 6]

END OF QUESTION PAPER

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The Periodic Table of the Elements

1 2

1	H	hydrogen	1
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relative atomic mass atomic symbol name atomic (proton) number

7	Li	lithium	3
9	Be	beryllium	4
11	Mg	magnesium	12
19	K	potassium	19
20	Ca	calcium	20
21	Sc	scandium	21
22	Ti	titanium	22
23	V	vanadium	23
24	Cr	chromium	24
25	Mn	manganese	25
26	Fe	iron	26
27	Co	cobalt	27
28	Ni	nickel	28
29	Cu	copper	29
30	Zn	zinc	30
31	Ga	gallium	31
32	Ge	germanium	32
33	As	arsenic	33
34	Se	selenium	34
35	Br	bromine	35
36	Kr	krypton	36
37	Rb	rubidium	37
38	Sr	strontium	38
39	Y	yttrium	39
40	Zr	zirconium	40
41	Nb	niobium	41
42	Mo	molybdenum	42
43	Tc	technetium	43
44	Ru	ruthenium	44
45	Rh	rhodium	45
46	Pd	palladium	46
47	Ag	silver	47
48	Cd	cadmium	48
49	In	indium	49
50	Sn	tin	50
51	Sb	antimony	51
52	Te	tellurium	52
53	I	iodine	53
54	Xe	xenon	54
55	Cs	caesium	55
56	Ba	barium	56
57	La*	lanthanum	57
58	Hf	hafnium	72
59	Ta	tantalum	73
60	W	tungsten	74
61	Re	rhenium	75
62	[261]	[261]	
63	Rf	rutherfordium	104
64	Bh	bohrium	107
65	Sg	seaborgium	106
66	Hs	hassium	108
67	Mt	meitnerium	109
68	Ds	darmstadtium	110
69	Rg	roentgenium	111

0	He	helium	2
1	H	hydrogen	1
2	C	carbon	6
3	N	nitrogen	7
4	O	oxygen	8
5	F	fluorine	9
6	Ne	neon	10
7	S	sulfur	16
8	Cl	chlorine	17
9	Ar	argon	18
10	Br	bromine	35
11	Ge	germanium	32
12	As	arsenic	33
13	Se	selenium	34
14	Ga	gallium	31
15	P	phosphorus	30
16	Al	aluminium	13
17	Si	silicon	14
18	B	boron	5
19	Li	lithium	3
20	Ca	calcium	20
21	Sc	scandium	21
22	Ti	titanium	22
23	V	vanadium	23
24	Cr	chromium	24
25	Mn	manganese	25
26	Fe	iron	26
27	Co	cobalt	27
28	Ni	nickel	28
29	Cu	copper	29
30	Zn	zinc	30
31	Ga	gallium	31
32	Ge	germanium	32
33	As	arsenic	33
34	Se	selenium	34
35	Br	bromine	35
36	Kr	krypton	36
37	Rb	rubidium	37
38	Sr	strontium	38
39	Y	yttrium	39
40	Zr	zirconium	40
41	Nb	niobium	41
42	Mo	molybdenum	42
43	Tc	technetium	43
44	Ru	ruthenium	44
45	Rh	rhodium	45
46	Pd	palladium	46
47	Ag	silver	47
48	Cd	cadmium	48
49	In	indium	49
50	Sn	tin	50
51	Sb	antimony	51
52	Te	tellurium	52
53	I	iodine	53
54	Xe	xenon	54
55	Cs	caesium	55
56	Ba	barium	56
57	La*	lanthanum	57
58	Hf	hafnium	72
59	Ta	tantalum	73
60	W	tungsten	74
61	Re	rhenium	75
62	[261]	[261]	
63	Rf	rutherfordium	104
64	Bh	bohrium	107
65	Sg	seaborgium	106
66	Hs	hassium	108
67	Mt	meitnerium	109
68	Ds	darmstadtium	110
69	Rg	roentgenium	111

* The lanthanoids (atomic numbers 58-71) and the actinoids (atomic numbers 90-103) have been omitted.

The relative atomic masses of copper and chlorine have not been rounded to the nearest whole number.

Elements with atomic numbers 112-116 have been reported but not fully authenticated